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 Worksheet Name: KEYChemistry 11 Stoichiometry Review Worksheet  
 Name: \_\_\_KEY\_\_\_ 33.  $\text{KClO}_4 \rightarrow 2\text{O}_2 + \text{KCl}$  Actual yield = 1.45g Theoretical yield = 5.95g  
 $\text{KClO}_4 \times 1 \text{ mol} = 122.55 \text{ g}$   $\times 2 \text{ O}_2 = 2 \times 32.00 \text{ g} = 64.00 \text{ g}$   
 $\text{KClO}_4 \times 32.00 \text{ g} = 1 \text{ mol} = 3.11 \text{ g O}_2$  % Yield =  $\frac{1.45 \text{ g}}{3.11 \text{ g}} \times 100\% = 46.6\%$   
 34.  $4\text{FeCO}_3 + \text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3 + 4\text{CO}_2$   
 Percent purity = 62.8%? g pure  $\text{FeCO}_3$  Jul 8th, 2022.  
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 Chlorate ( $\text{KClO}_3$ ) Is Heated, It Decomposes To Form Gases KCl And Oxygen ( $\text{O}_2$ .)  
 $2 \text{ KClO}_3 \rightarrow 2 \text{ KCl} + 3 \text{ O}_2$  C. Jan 3th, 2022.  
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